The influence of the additive BaGeO₃ on BaSnO₃ ceramics.

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Abstract. DTA, XRD and sintering investigations of the system BaSnO₃–BaGeO₃, prepared by a mixed-oxide method, are described herein. The melting temperature of this system is about 1270 ± 5 °C. We find a partial solubility of BaGeO₃ into BaSnO₃ of the order of 6–7 mol%. Up to 50 mol% BaGeO₃, the calcined powders (1150 °C) as well as the once-sintered samples consist of BaSnO₃ and orthorhombic BaGeO₃ at room temperature. A gradual appearance of hexagonal BaGeO₃ can be observed in calcined powders and once-sintered ceramics with a BaGeO₃ content above 50 mol%. After sintering at ≥ 1200 °C for more than 1 h all ceramic bodies consist of BaSnO₃ and orthorhombic BaGeO₃. The addition of BaGeO₃ leads to a considerable reduction of the

sintering temperature and to a strong densification. Sintering at 1180 °C for 10 h and an addition of only 1 mol% BaGeO₃ leads to dense ceramic bodies with cubic-like grains.

Keywords: Sintering (A); Perovskites (D); Powders-solid state reaction (A); Grain size (B); Ceramics

1. Introduction

Pure and doped barium stannate as well as its solid solutions (e.g. BaTi_{1-x}Sn_xO₃) have found important applications in materials science and technology due to their dielectric properties, semiconducting behaviours and high thermal stability. Because of these characteristic properties, BaSnO₃ based ceramics are becoming more and more important in material technology. It can be used to prepare thermally stable capacitors and to fabricate ceramic boundary layer capacitors [1,2,3,4,5,6,7,8]. Moreover, barium stannate can be also used as a functional material for semiconductor gas sensors [9,10,11,12,13,14,15] and photocatalytic applications [16,17,18]. In general, compacts on the basis of BaSnO₃ reveal only a moderate densification behaviour [2,19,20,21,22,23,24,25]. *Piercy* [26] obtained dense BaSnO₃ ceramic bodies at sintering temperatures above 1700 °C. Therefore, such ceramic bodies need high sintering temperatures or very long soaking times [7,27]. In order to produce capacitor components based on BaSnO₃, dense (almost pore free) material bodies are required, because pores would act as sink to the electrical charge carriers and would be the source of poor grain-to-grain connectivity and significant dielectric loss [23,24,28,29,30,31,32]. The addition of additives can significantly reduce the sintering temperature obtaining dense ceramic bodies due to an improvement of the densification behaviour. *Wang et al.* [33] used SiO₂ as a sintering aid for BaSnO₃ ceramics and *Kumar and Choudhary* [34] sintered BaSnO₃ at 1200 °C adding BaSiO₃. They found the

formation of solid solutions of the type $\text{BaSn}_{1-x}\text{Si}_x\text{O}_3$ ($x = 0\text{--}0.15$). The $\text{BaSn}_{1-x}\text{Si}_x\text{O}_3$ samples show NTCR (Negative Temperature Coefficient of Resistance) behaviour and a better electrical conduction at elevated temperature than pure BaSnO_3 ceramic bodies. BaGeO_3 can be also used as a sintering aid to drastically reduce the sintering temperature to produce dense ceramic bodies as reported for BaTiO_3 -based ceramics [35,36,37]. The influence and effect of the additive BaGeO_3 on the sintering behaviour and microstructure of BaSnO_3 ceramics or ceramic systems based on BaSnO_3 has not been investigated yet.

The aim of this work is to study the effect of BaGeO_3 on the sintering properties and microstructure of BaSnO_3 compacts. Furthermore, the formation of solid solutions and the phase evolution in the system $\text{BaSnO}_3\text{--BaGeO}_3$ is also investigated.

2. Experimental

2.1. Material preparation

$\text{BaSn}_{1-x}/\text{Ge}_x\text{O}_3$ powders with varying BaGeO_3 contents were prepared via a conventional mixed-oxide method. All powders were handled in a fume hood. Equivalent amounts of BaCO_3 (Saded VL 600, 99.9 %, *Solvay*), SnO_2 (≥ 99.0 %, *Merck*), and GeO_2 (99.999 %, *Acros Organics*) were milled in a PVC container for 24 h using ZrO_2 -balls and propan-2-ol ($m_{\text{powder}}:m_{\text{balls}}:m_{\text{Propan-2-ol}} = 1:1:4$). After filtering and drying the mixtures were calcined in static air (heating rate 10 K/min) for 2 h at 1150 °C.

For the shrinkage and sintering behaviour the calcined powders were milled with ZrO_2 -balls and propan-2-ol in a PVC container for 2 h ($m_{\text{powder}}:m_{\text{balls}} = 1:4$). After filtering and drying the powders were mixed with 5 mass% of a saturated aqueous solution of polyvinyl alcohol (PVA) as a pressing aid. Then the powders were pressed to pellets with a green density of about 3.4–3.6 g/cm³.

2.2. Analytical methods

X-ray powder diffraction (XRD) patterns were recorded by a *STOE STADI MP* diffractometer at 20 °C using $\text{CoK}\alpha_1$ radiation. The infrared spectra were recorded on a FTIR spectrometer *Mattson 5000 (Mattson Instruments Inc.)* with a resolution of 2 cm^{-1} as KBr pellets. Dilatometric investigations (shrinkage) were performed in a TMA 92-16.18 unit from *Setaram*. Differential thermoanalytic (DTA) measurements were recorded using a *Netzsch STA 449 System* (heating-/cooling rate 10 K/min, Pt crucible, flowing air (20 ml/min)). Each DTA sample was measured three times to ensure reproducibility. The melting temperature was determined by the onset temperature of the DTA signal [38].

The specific surface area was determined using nitrogen three-point BET (*Nova 1000, Quantachrome Corporation*). The equivalent BET particle diameter was calculated according to the equation given in ref. [39] assuming the powder particles were spherical or cubic in shape. Scanning electron microscope images and energy dispersive X-ray analyses (EDX) were recorded with a *Philips XL30 ESEM* (Environmental Scanning Electron Microscope) and an attached energy dispersive X-ray spectrometer from *Edax*. Transmission electron microscope images were recorded with a *Jeol JEM 2100F*.

3. Results and discussion

3.1. Characterization of the system $\text{BaSnO}_3\text{--BaGeO}_3$

Mixtures of BaCO_3 , SnO_2 and GeO_2 were heated in a Pt-crucible at 1150 °C for 2 h to obtain powders with a composition of $\text{BaSn}_{1-x}\text{Ge}_x\text{O}_3$. Calcined powders with 1 and 2 mol% BaGeO_3 show only reflections of the BaSnO_3 phase (perovskite structure) (Fig. 1). A powder with 5 mol% BaGeO_3 hints at very small traces of an additional

orthorhombic BaGeO₃ phase. Up to a BaGeO₃ content less than 50 mol% we observe always reflections of BaSnO₃ and orthorhombic BaGeO₃. BaGeO₃ crystallises in two modifications. Hexagonal (α)-BaGeO₃ (low-temperature form) having a pseudowollastonite-type structure [40] and orthorhombic (β)-BaGeO₃ (high-temperature form) crystallises in a pyroxene-type structure [41]. The phase transition between these modifications appears at about 1200 °C [42]. At 50 mol% BaGeO₃ we obtain both the orthorhombic BaGeO₃ phase and traces of the hexagonal one. The fraction of the hexagonal phase raises with increasing BaGeO₃ content. Fig. 2 shows IR spectra of calcined powders with 1 and 15 mol% BaGeO₃, exemplarily. The peak around 1420 cm⁻¹ suggests very small traces of BaCO₃ [43].

The above described calcined powders were the basic powders for all further investigations. After pelleting of the these powders, sintering at 1150 °C for 10 h and cooling down (10 K/min) to room temperature, ceramics with a BaGeO₃ content $\geq$ 50 mol% show besides the BaSnO₃ phase only reflections of hexagonal BaGeO₃ (Fig. 3). In contrast, samples with less than 50 mol% BaGeO₃ reveal the pattern of orthorhombic BaGeO₃. However, after repeated crushing and sintering at 1150 °C, the XRD patterns of these samples (< 50 mol% BaGeO₃) reveal the formation of hexagonal BaGeO₃ (Fig. 3e). After sintering at 1200 °C for 10 h we observe that all ceramic samples exhibit only reflections of orthorhombic BaGeO₃. Up to 1400 °C we found as crystalline phases always both BaSnO₃ and BaGeO₃. It can be seen from Fig. 3b, that the BaGeO₃ phase clearly appears at samples with at least 10 mol% BaGeO₃ content. Sintered bodies with 5 mol% BaGeO₃ hint at very small traces of a BaGeO₃ phase. However, repeated sintering, crushing and pelleting of this sample leads to a disappearing of the BaGeO₃ reflections and thus to a complete formation of a solid solution between the non-isotypic compounds BaSnO₃ and BaGeO₃.

For the following DTA and EDX measurements the calcined samples were pressed to pellets, sintered several times with intermittent cooling, crushing, and re-pelleting. EDX measurements of several grains in various ceramic bodies suggest a solid solubility of BaGeO₃ into the BaSnO₃ structure of about 6–7 mol%, which confirm the above XRD results. Because of the limited solid solubility, we describe the system BaSnO₃–BaGeO₃ by the formula BaSn_{1-x}/Ge_xO₃.

DTA investigations of pure BaGeO₃ indicated a melting point of 1290 ± 5 °C (Fig. 4) and a hexagonal $\rightleftharpoons$ orthorhombic phase transition temperature of 1190 ± 5 °C (not shown in Fig. 4) in good agreement with earlier data [42,44]. BaSnO₃ has a very high melting point, which could not be determined by us. However, *Wagner and Binder* [45] reported a melting temperature of 2057 °C. DTA measurements of BaSn_{1-x}/Ge_xO₃ samples with various BaGeO₃ contents ($x \geq 0.1$) reveal a melting temperature at about 1270 ± 5 °C (Fig. 4). Samples with a BaGeO₃ concentration below 10 mol% do not show any melting peak.

3.2. Shrinkage and sintering behaviour

Four different preceramic powders were used for these investigations. Powders with BaGeO₃ contents of 1 mol% (**1**), 2 mol% (**2**), 5 mol% (**5**), and 15 mol% (**15**) were obtained after calcination at 1150 °C for 2 h (heating rate 10 K/min), as described above. The specific surface areas and the equivalent BET particle diameters of these powders are given in Tab. 1. It can be seen that an increasing BaGeO₃ content causes a decreasing specific surface area and thus a raising BET particle diameter, which can be considered as an average size of the primary particles [46]. Fig. 5 shows the cubic- and cuboid-like morphology of powder **1**. The particles of this powder are strong agglomerated (Fig. 5b). The theoretical bulk densities of the resulting ceramic bodies

were calculated as 7.22 g/cm³ (**1**), 7.19 g/cm³ (**2**), 7.12 g/cm³ (**5**), and 6.85 g/cm³ (**15**) [47].

The shrinkage behaviour of green bodies of **1**, **5**, and **15** are depicted in Fig. 6. Sample **1** slowly starts to shrink at about 1050 °C and the maximum shrinkage rate is observed at 1295 °C (-2.3 %/min). Further addition of BaGeO₃ leads to a shift of the maximum shrinkage rate to 1265 °C (-1.8 %/min) (**5**). The maximum shrinkage rate in sample **15** is characterised by a sharp asymmetric peak. The peak at 1280 °C exhibits an approximately two times higher shrinkage rate (-3.8 %/min) and is assigned to the formation of a liquid phase. In samples with a BaGeO₃ content > 15 mol% we observe the appearance of two separate peaks; a first peak in the range 1200–1215 °C and a second one at about 1280 °C. The shrinkage of sample **5** slowly starts at 950 °C, however a significant shrinkage occurs above 1060 °C, whereas the sample containing 15 mol% BaGeO₃ (**15**) starts to shrink not until about 1040 °C. Compared to pure BaSnO₃ samples [19] the addition of BaGeO₃ causes considerably higher shrinkage rates and thus stronger shrinkage at lower temperatures.

Fig. 7 shows the final bulk densities of BaSn_{1-x}/Ge_xO₃ ($x = 0.01$ (**1**); 0.02 (**2**); 0.05 (**5**); 0.15 (**15**)) ceramics after isothermal sintering between 1150 °C and 1400 °C for 1 and 10 h, respectively. Up to 1200 °C and a soaking time of 1 h ceramic bodies with a relative density of 76–87 % can be obtained. Dense ceramic bodies (rel. density ≥ 90 %) can be obtained at temperatures above 1200 °C. As seen in Fig. 7, after sintering of 1 h ceramics of **1**, **2** and **5** achieve at 1230 °C relative densities of 92 % (6.62 g/cm³), 91 % (6.53 g/cm³) and 93 % (6.65 g/cm³), respectively. Sample **15** requires a sintering temperature of 1250 °C to form dense ceramic bodies. Higher sintering temperatures lead only to a marginal increase in density. A prolonged soaking time of 10 h results in a slight reduction of the required sintering temperature for dense ceramic bodies. After 10 h and a temperature of 1180 °C we obtain ceramics with densities of 6.52 g/cm³ (90

%, **1**), 6.52 g/cm³ (91 %, **2**), 6.44 g/cm³ (90 %, **5**) and 5.58 g/cm³ (81 %, **15**). Above 1150 °C the final relative densities of ceramics **1**, **2** and **5** differs only slightly from each other, whereas ceramic bodies of **15** exhibit lower relative densities. The general decrease in absolute densities of BaSn_{1-x}/Ge_xO₃ ceramics with increasing BaGeO₃ content is also due to the reduction of the maximum achievable density by adding BaGeO₃ (see above).

As seen in Fig. 3c ceramic bodies of **15** consist of orthorhombic BaGeO₃ besides BaSnO₃ as the main crystalline phase.

From the outlined results it can be concluded that an amount of 1 mol% BaGeO₃ is sufficient to obtain dense BaSn_{1-x}/Ge_xO₃ ceramic bodies at low sintering temperatures.

The micrographs of ceramics of **1** are shown in Fig. 8. Sintering at 1230 °C for 1 h of **1** results in ceramic bodies with cubic-like grains of about 0.3–1.3 µm ($\varnothing = 0.7 \mu\text{m}$ [48]), whereas ceramic bodies of **15** (not shown in Fig. 8) reveal grain sizes between 0.5 and 1.8 µm. A sintering temperature of 1300 °C for 1 h (Fig. 8b) reveal in both samples a bimodal or heterogeneous grain size distribution with grain sizes of about 0.5–1 µm/ 2.5–7 µm (**1**) and 0.7–6 µm (**15**). Sintering at 1400 °C (Fig. 8c) leads also to a bimodal microstructure with irregular grains between 2.5–26 µm (**1**). An increasing soaking time to 10 h at 1180 °C causes for ceramics of **1** a homogeneous microstructure with grains between 0.4–2 µm ($\varnothing = 1.6 \mu\text{m}$) (Fig. 8d) and in sample **15** we find grains in the range 0.8–2.8 µm. Sintering at 1250 °C for 10 h (Fig. 8e) leads in both samples (**1**, **15**) to a bimodal grain growth with grain fractions of 0.8–2 µm and 5–20 µm (**1**) or to a heterogeneous microstructure with grains between 1–15 µm (**15**). The smaller grain fractions mainly consist of cubic-like grains, whereas the larger grains have a more irregular shape as seen from the inset in Fig 8e.

The isothermal sintering clearly shows, that dense ceramic bodies can be obtained at sintering temperatures below the formation of the liquid phase ($< 1270\text{ }^{\circ}\text{C}$). This indicated that the formation of a liquid phase is not essential for densification of the investigated ceramic bodies. Therefore, the additive BaGeO_3 does not affect primarily as a liquid phase former, but it improves the sintering behaviour by sliding processes (viscous flow). Fig. 9 exemplarily shows the result of isothermal dilatometric investigations of compacts of **2** and **15**. By means of the general shrinkage equation **I** [49], a logarithmic representation of $\Delta L/L_0$ versus time (t) allows to determine the dominant shrinkage mechanisms.

$$\left(\frac{\Delta L}{L_0}\right)^{\frac{m}{2}} = -\frac{H}{2^m R^n} t \quad (\text{I})$$

Here, $\Delta L/L_0$ is the relative shrinkage, H is a function containing material parameters, R is the radius of the particles, t is the time, n and m are numerical exponents depending on the shrinkage mechanism.

The isothermal part clearly shows that the isothermal shrinkage process is divided into two sections representing two different shrinkage mechanisms. [35,50]. The first segment of the isothermal curves has a slope according to a shrinkage exponent of about $m = 2$ indicating viscous flow as the major shrinkage mechanism. The second segment ($m \approx 4\text{--}6$) suggests grain boundary diffusion and volume diffusion from the grain boundaries as the dominant shrinkage process. It can be seen that the beginning of the shrinkage process is dominated by sliding processes (viscous flow). Such processes are caused by defect-rich and amorphous (glass-like) contact boundaries, respectively [51,52,53]. The sliding processes of whole grains are supported by BaGeO_3 and result in a reducing sintering temperature. A similar improved densification behaviour during solid-state sintering by addition of germanium- and silicon-based additives is also observed in BaTiO_3 ceramics [35,50].

4. Conclusion

The system $\text{BaSnO}_3\text{--BaGeO}_3$ ($\text{BaSn}_{1-x}\text{Ge}_x\text{O}_3$, $x \geq 0.1$) reveals a melting temperature of about 1270 ± 5 °C. XRD investigations and EDX measurements of single grains of various ceramic bodies suggest a solubility of BaGeO_3 in BaSnO_3 of the order of 6–7 mol%. $\text{BaSn}_{1-x}\text{Ge}_x\text{O}_3$ samples ($x \geq 0.1$) heated up to 1400 °C show reflections of both a BaSnO_3 phase and a BaGeO_3 one, whereas samples with $x < 0.1$ reveal only reflections of the BaSnO_3 phase.

BaGeO_3 can be used as a sintering additive for BaSnO_3 ceramics to reduce the sintering temperature. The ceramic bodies were prepared from mixed-oxid powders (calcined at 1150 °C, 2 h). The sintering behaviour was studied on samples containing 1–15 mol% BaGeO_3 . An addition of 1 mol% BaGeO_3 is sufficient to obtain dense ceramic bodies (relative density ≥ 90 %) after a conventional solid state sintering process of 1 h at 1230 °C, or 10 h at 1180 °C. Raising sintering temperatures or an increased BaGeO_3 content promote the formation of an heterogeneous or bimodal grain size distribution in the ceramic bodies.

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Table 1

Properties of $\text{BaSn}_{1-x}/\text{Ge}_x\text{O}_3$ powders calcined at 1150 °C for 2 h.

Powder	BaGeO_3 content (mol%)	BET specific surface area (m^2/g)	Average particle diameter* (nm)
1	1	3.9	213
2	2	2.7	309
5	5	2.4	351
15	15	1.5	581

* calculated from the specific surface area [39]

Fig. 1. XRD patterns (recorded at 20 °C) of calcined $\text{BaSn}_{1-x}/\text{Ge}_x\text{O}_3$ powders (1150 °C, 2 h) with various BaGeO_3 contents. (a) 1 mol%; (b) 5 mol%; (c) 10 mol%; (d) 15 mol%; (e) 30 mol%; (f) 50 mol%; (g) 80 mol%.

Fig. 2. IR spectra of calcined samples containing (a) 1 mol% and (b) 15 mol% BaGeO₃. Spectrum (a) shows a weak broad band at about 1630 cm⁻¹ (δ_{OH}) due to water absorbed by the KBr powder.

Fig. 3. XRD patterns (recorded at 20 °C) of $\text{BaSn}_{1-x}\text{Ge}_x\text{O}_3$ ceramics after various sintering temperatures and soaking times. (a–d) 1150 °C, 10 h; (e) 1150 °C, 10 h with intermittent cooling, crushing and pelleting; (f, g) 1150 °C, 10 h; (h, i) 1200 °C, 10 h.

Fig. 4. DTA traces (melting point) of BaGeO_3 and various $\text{BaSn}_{1-x}\text{Ge}_x\text{O}_3$ ($x \geq 0.15$) samples.

Fig. 5. TEM images of powder 1.

Fig. 6. Shrinkage behaviour (non-isothermal, heating rate 10 K/min) of green bodies of 1, 5 and 15. The inset shows the relative shrinkage rate ($d(\Delta L/L_0)/dt$) of these bodies.

Fig. 7. Final densities of $\text{BaSn}_{1-x}\text{Ge}_x\text{O}_3$ ceramic bodies of 1 ($x=0.01$), 2 ($x=0.02$), 5 ($x=0.05$) and 15 ($x=0.15$) after an isothermal sintering process at various temperatures at the indicated soaking times in a muffle furnace (heating/cooling rate: 10 K/min). (a) Relative densities vs. sintering temperature and (b) absolute densities vs. sintering temperature. The relative densities are related to the theoretical density of 7.22 g/cm³ (1), 7.19 g/cm³ (2), 7.12 g/cm³ (5) and 6.85 g/cm³ (15).

Fig. 8. SEM micrographs of the surfaces of ceramic bodies of 1 after various sintering temperatures and soaking times. (a) 1230 °C, 1 h; (b) 1300 °C, 1 h; (c) 1400 °C, 1 h; (d) 1180 °C, 10 h; (e) 1250 °C, 10 h.

Fig. 9. Isothermal dilatometric investigations demonstrated at 1150 °C for green bodies of 2 and 15. The figure shows a double logarithmic plot of $\Delta L/L_0$ vs. time (t) of the isothermal segments (the initial values were set to zero).

References

- [1] R. Vivekanandan and T. R. N. Kutty, Grain boundary layer ceramic capacitors based on donor-doped $\text{Ba}(\text{Ti}_{1-x}\text{Sn}_x)\text{O}_3$. *Mater. Sci. Eng. B*, 1990, **6**[4], 221–231.
- [2] P. Singh, B. J. Brandenburg, C. P. Sebastian, P. Singh, S. Singh, D. Kumar, and O. Parkash, Electronic structure, electrical and dielectric properties of BaSnO_3 below 300 K. *Jpn. J. Appl. Phys.*, 2008, **47**, 3540–3545
- [3] A. Movchikova, O. Malyshkina, G. Suchanek, G. Gerlach, R. Steinhausen, H. T. Langhammer, C. Pientzsche, and H. Beige, Study of the pyroelectric behavior of $\text{BaTi}_{1-x}\text{Sn}_x\text{O}_3$ piezo-ceramics. *J. Electroceram.*, 2008, **20**, 43–46.
- [4] T. Wang, X. M. Chen, and X. H. Zheng, Dielectric characteristics and tunability of barium stannate titanate ceramics. *J. Electroceram.*, 2003, **11**, 173–178
- [5] A. Kumar, B. P. Singh, R. N. P. Choudhary, and A. K. Thakur, Ferroelectric phase transition in Te-modified BaSnO_3 . *Materials Letters*, 2005, **59**, 1880–1888.
- [6] O. I. Prokopalo, Conductivity and anomalous polarization in ceramic ferroelectrics with perovskite structure. *Ferroelectrics*, 1976, **14**, 683–685.
- [7] P. Singh, D. Kumar, and O. Parkash, Dielectric behavior of the system $\text{BaSn}_{1-x}\text{Nb}_x\text{O}_3$ ($x \leq 0.10$). *J. Appl. Phys.*, 2005, **97**, 074103pp.
- [8] H. Brauer, Korngrenzensperrschichten in BaTiO_3 Keramik mit hoher effektiver Dielektrizitätskonstante. *Z. Angew. Phys.*, 1970, **29**, 282–287.

-
- [9] B. Ostrick, M. Fleischer, U. Lampe, and H. Meixner, Preparation of stoichiometric barium stannate thin films: Hall measurements and gas sensitivities. *Sens. Actuators B*, 1997, **44**, 601–606.
- [10] S. Tao, F. Gao, X. Liu, and O. T. Sørensen, Ethanol-sensing characteristics of barium stannate prepared by chemical precipitation. *Sens. Actuators B*, 2000, **71**, 223–227.
- [11] S. Hodjati, K. Vaezzadeh, C. Petit, V. Pitchon, and A. Kiennemann, NO_x sorption–desorption study: application to diesel and lean-burn exhaust gas (selective NO_x recirculation technique). *Catal. Today*, 2000, **59**, 323–334.
- [12] P. McGeehin and D. E. Williams, Sensing gaseous substances. Int. Appl., Patent No. WO 9308467 A1 19930429, 1993.
- [13] V. Gopal Reddy, S. V. Manorama, and V. J. Rao, Preparation and characterization of barium stannate: application as a liquefied petroleum gas sensor. *J. Mater. Sci.:Mater. Electron.*, 2001, **12**, 137–142.
- [14] S. Upadhyay, and P. Kavitha, Lanthanum doped barium stannate for humidity sensor. *Materials Letters*, 2007, **61**, 1912–1915.
- [15] J. Cerda, J. Arbiol, G. Dezanneau, R. Diaz, and J. R. Morante, Perovskite-type BaSnO₃ powders for high temperature gas sensor applications. *Sens. Actuators B*, 2002, **84**, 21–25.
- [16] P. H. Borse, U. A. Joshi, S. M. Ji, J. S. Jang, J. S. Lee, E. D. Jeong, and H. G. Kim, Band gap tuning of lead-substituted BaSnO₃ for visible light photocatalysis. *Appl. Phys. Lett.*, 2007, **90**, 034103–1pp.
- [17] Y. Yuan, J. Lv, X. Jiang, Z. Li, T. Yu, Z. Zou, and J. Ye, Large impact of strontium substitution on photocatalytic water splitting activity of BaSnO₃. *Appl. Phys. Lett.*, 2007, **91**, 094107–1pp.

-
- [18] Y. Zhang, H. Zhang, Y. Wang, and W. F. Zhang, Efficient visible spectrum sensitization of BaSnO₃ nanoparticles with N719. *J. Phys. Chem. C*, 2008, **112**, 8553–8557.
- [19] R. Köferstein, L. Jäger, M. Zenkner, and S. G. Ebbinghaus, Nano-sized BaSnO₃ powder via a precursor route. Comparative study of sintering behaviour and mechanism of fine and coarse-grained powders. *J. Eur. Ceram. Soc.*, 2009, **29**, 2317–2324.
- [20] P. Singh, D. Kumar, and O. Parkash, DC conduction behaviour of niobium doped barium stannate. *J. Mater. Sci.: Mater. Electron.*, 2005, **16**, 145–148.
- [21] S. Upadhyay, O. Parkash, and D. Kumar, Dielectric relaxion and conduction in the system Ba_{1-x}La_xSn_{1-x}Cr_xO₃. *J. Mater. Sci.:Mater. Electron.*, 2001, **12**, 165–172.
- [22] X. Wie and X. Yao, Preparation, structure and dielectric property of barium stannate titanate ceramics. *Mater. Sci. Eng. B*, 2007, **137**, 184–188.
- [23] M. Cernea, A. Manea, D. Piazza, C. Galassi, and E. Vasile, Ba(Ti_{1-x}Sn_x)O₃ (x = 0.13) dielectric ceramics prepared by coprecipitation. *J. Am. Ceram. Soc.* 2007, **90**, 1728–1732.
- [24] A.-M. Azad and N. C. Hon, Characterization of BaSnO₃-based ceramics. Part 1. Synthesis, processing and microstructural development. *J. Alloys Comp.*, 1998, **270**, 95–106.
- [25] L.-C. Tien, C.-C. Chou, and D.-S. Tsai, Microstructure of Ba(Mg_{1/3}Ta_{2/3})O₃-BaSnO₃ microwave dielectrics. *Ceram. Inter.*, 2000, **26**, 57–62.
- [26] B. Piercy, Dielectric properties of some polycrystalline stannates and cerates. *Trans. Faraday Soc.*, 1959, **55**, 39–51.

-
- [27] J. Cava, P. Gammel, B. Batlogg, J. J. Krajewski, W. F. Peck, Jr., W. L. Rupp, Jr., R. Felder, and R. B. van Dover, Nonsuperconducting $\text{BaSn}_{1-x}\text{Sb}_x\text{O}_3$: The 5s-orbital analog of $\text{BaPb}_{1-x}\text{Bi}_x\text{O}_3$. *Phys. Rev. B*, 1990, **42**, 4815–4818.
 - [28] R. W. Rice, *Porosity of ceramics*, ed. Marcel Dekker, New York, 1998, pp. 325 et seqq.
 - [29] S. Marković, M. Mitrić, N. Cvjetičanin, and D. Uskoković, Structural and dielectric properties of $\text{BaTi}_{1-x}\text{Sn}_x\text{O}_3$ ceramics. *Mater. Sci. Forum*, 2006, **518**, 241–246.
 - [30] P. Singh, O. Parkash, and D. Kumar, Electrical conduction behaviour of perovskite oxide system $\text{BaSn}_{1-x}\text{Nb}_x\text{O}_3$. *Solid State Ionics*, 2005, **176**, 2167–2170.
 - [31] Z. He, J. Ma, and R. Zhang, Investigation on the microstructure and ferroelectric properties of porous PZT ceramics. *Ceram. Inter.*, 2004, **30**, 1353–1356.
 - [32] D. M. Iddles, A. J. Bell, and A. J. Moulson, Relationships between dopants, microstructure and the microwave dielectric properties of $\text{ZrO}_2\text{-TiO}_2\text{-SnO}_2$ ceramics. *J. Mater. Sci.*, 1992, **27**, 6303–6310.
 - [33] Z. Wang, F. Zhou, and Z. Chen, Preparation and electric properties of BaSnO_3 ceramics. *Dianzi Yuanjian Yu Cailiao*, 2006, **25**, 58–60.
 - [34] A. Kumar and R. N. P. Choudhary, Characterization of electrical behaviour of Si modified BaSnO_3 electroceramics using impedance analysis. *J. Mater. Sci.*, 2007, **42**, 2476–2485.
 - [35] R. Köferstein, L. Jäger, M. Zenkner, T. Müller, and H.-P. Abicht, Shrinkage mechanism and phase evolution of fine-grain BaTiO_3 powder compacts containing 10 mol% BaGeO_3 prepared via a precursor route. *Mater. Chem. Phys.*, 2008, **112**, 531–535.

-
- [36] M. Zenkner, L. Jäger, R. Köferstein, and H.-P. Abicht, Synthesis and characterization of nanoscaled $\text{Ba}(\text{Ti}_{1-x-y}\text{Sn}_x\text{Ge}_y)\text{O}_3$ powders and corresponding ceramics. *Solid State Sci.*, 2008, **10**, 1556–1562.
- [37] R. Köferstein, L. Jäger, M. Zenkner, and H.-P. Abicht, Preparation and sintering behaviour of a fine grain BaTiO_3 powder containing 10 mol% BaGeO_3 . *J. Mater. Sci.*, 2008, **43**, 832–838.
- [38] H. K. Cammenga and M. Epple, Grundlagen der Thermischen Analysetechniken und ihre Anwendungen in der präparativen Chemie. *Angew. Chem.*, 1995, **107**, 1284–1301.
- [39] V. D. Allred, S. R. Buxton, and J. P. McBride, Characteristic properties of thorium oxide particles. *J. Phys. Chem.*, 1957, **61**, 117–120.
- [40] F. Liebau, Zur Kristallchemie der Silikate, Germanate, Fluoberyllate des Formeltyps ABX_3 . *Neues Jahrb. Mineral.*, 1960, **94**, 1209–1222.
- [41] W. Hilmar, Die Struktur der Hochtemperaturform des Bariumgermanates BaGeO_3 (h). *Acta Cryst.*, 1962, **15**, 1101–1105.
- [42] J. P. Guha, D. Kolar, and A. Porenta, Polymorphic modifications of barium metagermanate. *J. Thermal. Anal.*, 1976, **9**, 37–41.
- [43] K. Nakamoto, *Infrared and raman spectra of inorganic and coordination compounds*, ed. John Wiley & Sons, USA, 1979, p. 129.
- [44] J. P. Guha and D. Kolar, Phase equilibria in the system BaTiO_3 – BaGeO_3 . *J. Mater. Sci.*, 1972, **7**, 1192–1196.
- [45] G. Wagner and H. Binder, Untersuchung der binären Systeme BaO – SnO_2 und BaO – PbO_2 . I. Phasenanalysen. *Z. Allg. Anorg. Chem.*, 1958, **297**, 328–346.

-
- [46] M. T. Buscaglia, M. Bassoli, V. Buscaglia and R. Alessio, Solid-State synthesis of ultrafine BaTiO₃ powders from nanocrystalline BaCO₃ and TiO₂, *J. Am. Ceram. Soc.*, 2005, **88** [9] , 2374–2379.
- [47] G. W. Marks and L. A. Monson, Effect of certain group IV oxides on dielectric constant and dissipation factor of barium titanate. *Ind. Eng. Chem.*, 1955, **47**(8), 1611–1620.
- [48] J. C. Wurst, and J. A. Nelson, Lineal intercept technique for measuring grain size in two-phase polycrystalline ceramics. *J. Am. Ceram. Soc.*, 1972, **55**, 109.
- [49] M. N. Rahaman, *Ceramic proceeding and sintering*, ed. Marcel Dekker, New York, 1995, pp. 398 et seqq.
- [50] D. Völtzke and H.-P. Abicht, The influence of different additives and the mode of their addition on the sintering behaviour and the properties of semiconducting barium titanate ceramics. *Solid State Sci.*, 2000, **2**, 149–159.
- [51] Ya. E. Geguzin and Yu. I. Klinchuk, Mechanism and kinetics of the initial stage of the solid-phase sintering of pressed parts from crystal-body powders activity during sintering. *Poroshkovaya Metallurgiya*, 1976, **7**, 17–25.
- [52] W. Schatt, Vorgänge beim Festphasensintern und ihre Verallgemeinerungsfähigkeit. *Z. Metallkde.*, 1989, **80** [11], 809–816.
- [53] W. Schatt, Study of sintering processes by positron annihilation. *Solid State Phenomena*, 1992, **25–26**, 23–28.